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products & reactants-Mass to Mole Conversions-Multiply the grams by the g/mol ratio in the reference table-Ex:  $15\text{g of C} \times \frac{1}{12.011} = 1.3 \text{ grams}$ -Molar Ratio-Find the ratio of between 2 substances in a formula-Ex:  $\text{CuCl} + \text{Zn} \rightarrow \text{ZnCl} + \text{Cu}$ -The molar mass formula between Cu and Zn is 1 to 1-GFM-The mass in grams of 1 mole of any substance ' s ...

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1. Define the following: a. Stoichiometry-the study of the quantitative relationships between the amounts of reactants used and the products formed by a chemical



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In a typical Stoichiometry problem, the given quantity is first converted to \_\_\_\_\_. Then, the mole ratio from the balanced equation is used to calculate the moles of unwanted substance. ... Chem Quiz 12.1-12.3. 20 terms ...

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, 154.21 g/mol ( $6 \cdot 12 + 5 \cdot 1 = 77$ :  $154/77 = 2$ ) C<sub>12</sub> H<sub>10</sub> C<sub>3</sub> H<sub>6</sub> O<sub>2</sub>, 222.24 g/mol ( $12 \cdot 3 + 6 \cdot 1 + 2 \cdot 16 = 77$ :  $222/74 = 3$ ) C<sub>9</sub> H<sub>18</sub> O<sub>6</sub> 3. Empirical Formula from Percentage Composition (don't round until the very end!!!!) A compound is composed of 12.590% hydrogen, 37.404% aluminum, and 50.006% carbon. Please find the empirical formula:

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nonchemical process. It teaches you how to interpret chemical equations in terms of interacting moles, representative particles, masses, and

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equation that follows each problem, write in the space

provided the mole ratio that can be used to solve the

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