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$\ln K = (\Delta_r S/R) - (\Delta_r H/RT)$ The equilibrium for the dissociation of one mole of iodine vapour is 3.94×10^{-4} at 900K and the standard enthalpy is 154 kJmol⁻¹. Calculate the equilibrium constant at 1200K assuming

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the standard enthalpy remains constant. $\ln (K_2/K_1) = - \Delta_r H/R (1/T_2 - 1/T_1)$, ANS = 6.7×10^{-2} at 1200K.

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Intended for students of intermediate organic chemistry, this text shows how to write a reasonable mechanism for an organic chemical transformation. The discussion is organized by types of mechanisms and the conditions under which the reaction is executed, rather than by the overall reaction as is the case in most textbooks. Each chapter discusses common mechanistic pathways and suggests practical tips for drawing them. Worked problems are included in the discussion of each

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mechanism, and "common error alerts" are scattered throughout the text to warn readers about pitfalls and misconceptions that bedevil students. Each chapter is capped by a large problem set.

Following in the tradition of the first four editions, the goal of this market leading textbook, "Chemistry in Context," fifth edition, is to establish chemical principles on a need-to-know basis within a contextual framework of significant social, political, economic and ethical issues. The non traditional approach of "Chemistry in Context" reflect today's technological issues and the chemistry principles imbedded within them. Global warming, alternate fuels, nutrition, and genetic engineering are examples of issues that are covered in CIC.

Organic chemistry courses are often difficult for students, and instructors are constantly seeking new ways to improve student learning. This volume details active learning strategies implemented at a variety of institutional settings, including small and large; private and public; liberal arts and technical; and highly selective and open-enrollment institutions. Readers will find detailed descriptions of methods and materials, in addition to data supporting

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analyses of the effectiveness of reported pedagogies.

The series "Handbook of Green Chemistry" edited by P. Anastas who is one leading pioneer of this field is the ultimate reference. Volume 10 of the Handbook of Green Chemistry presents important tools, databases, and laboratory approaches to support chemists in academia and industry to achieve their green chemistry goals.

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