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Answer the following questions in the space provided. 1. Given the following equation: $C_3H_4(g) + xO_2(g) \rightarrow 3CO_2(g) + 2H_2O(g)$

a. What is the value of the coefficient x in this equation? 40.07 g/mol

b. What is the molar mass of C_3H_4 ? 2 mol O_2 :1 mol H_2O

c. What is the mole ratio ...

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Modern Chemistry Section Review
Stoichiometry Answers Mix & match
(both simple stoichiometry and limiting reagent problems) Units to use (select at least one): Grams Moles Particles (e.g. atoms/molecules/formula units) Chemical formulas or names: Formulas only Names only Mix & match formulas &

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Stoichiometry CHAPTER 9 REVIEW
Answers
Stoichiometry SECTION 3 PROBLEMS

Write the answer on the line to the left.

Show all your work in the space provided.

1. _____ The actual yield of a reaction is 22 g and the theoretical yield is 25 g.

Calculate the percentage yield. 2. 6.0 mol

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Stoichiometry. composition stoichiometry deals with the mass relationships of elements in compounds. reaction stoichiometry involves the mass relationships between reactants and products in a chemical reaction.

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The Stoichiometry chapter of this Holt McDougal Modern Chemistry Companion Course helps students learn the essential chemistry lessons of stoichiometry. Holt McDougal Modern Chemistry Chapter 9: Stoichiometry ...

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Modern Chemistry □ CHAPTER 9

HOMEWORK 9-2 (pp. 280-282)

VOCABULARY Complete each sentence.

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more sensitively some of the offerings of
the standard text book of general

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chemistry. As a supplement thereto, it covers various aspects of formulation and stoichiometry that are frequently treated far too perfunctorily or, in many instances, are not considered at all. The inadequate attention often accorded by the comprehensive text to many topics within its proper purview arises, understandably enough, from the numerous broad and highly varied objectives set for the first year of the curriculum for modern chemistry in colleges and universities. For the serious student this means, more often than not, the frustrations of questions unanswered. The amplification that this book proffers in the immediate area of its subject covers the equations representing internal redox reactions, not only of the simple but, also, of the multiple disproportionations of which the complexities often discourage an undertaking despite the challenge they

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offer: distinctions to be observed in the balancing of equations in contrasting alkali-basic and ammonia-basic reaction media; quantitative contributions made by the ionization or dissociation effects of electrolytes to the colligative properties of their solutions; intensive application of the universal reaction principle of chemical equivalence to the stoichiometry of oxidation and reduction.

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Fundamentals of Chemistry, Fourth Edition covers the fundamentals of chemistry. The book describes the formation of ionic and covalent bonds; the Lewis theory of bonding; resonance; and the shape of molecules. The book then discusses the theory and some applications of the four kinds of spectroscopy: ultraviolet, infrared, nuclear (proton) magnetic resonance, and mass. Topics that combine environmental significance with descriptive chemistry, including atmospheric pollution from automobile exhaust; the metallurgy of iron and aluminum; corrosion; reactions involving ozone in the upper atmosphere; and the methods of controlling the pollution of air and water, are also considered. Chemists and students taking courses related to chemistry and environmental chemistry will find the book invaluable.

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Long considered the standard for honors and high-level mainstream general chemistry courses, PRINCIPLES OF MODERN CHEMISTRY continues to set the standard as the most modern, rigorous, and chemically and mathematically accurate text on the market. This authoritative text features an atoms first approach and thoroughly revised chapters on Quantum Mechanics and Molecular Structure (Chapter 6), Electrochemistry (Chapter 17), and Molecular Spectroscopy and Photochemistry (Chapter 20). In addition, the text utilizes mathematically accurate and artistic atomic and molecular orbital art, and is student friendly without compromising its rigor. End-of-chapter study aids now focus on only the most important key objectives, equations and concepts, making it easier for students to locate chapter content, while new

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applications to a wide range of disciplines, such as biology, chemical engineering, biochemistry, and medicine deepen students' understanding of the relevance of chemistry beyond the classroom.

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